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In the 4th Edition the first law of thermodynamics is presented in a single chapter, using a highly intuitive and unified approach. Over 200 multiple-choice problems at the end of chapters prepare the students for the Fundamentals of Engineering (FE) exam, and can also be used for general review and quizzing.

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Multiply molarity by the solution volume (in liters) and the number 10<sup>6</sup>—the molar mass of sodium carbonate—to calculate the mass of sodium carbonate needed. For example, to make 300 mL of 0.2 molar solution, you will need:  $0.2 \times 0.3 \text{ L} \times 10^6 = 6.36 \text{ g}$  Note that 300 mL = 0.3 L

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This text looks at thermodynamics and statistical mechanics. Part I introduces concepts of thermodynamics and statistical mechanics from a unified view. ... Fundamentals of Statistical and Thermal Physics. ... places in explaining the derivations of key ideas step by step more than makes up for the lack of an easily accessible solution manual ...

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The ideal gas law, also called the general gas equation, is the equation of state of a hypothetical ideal gas. It is a good approximation of the behavior of many gases under many conditions, although it has several limitations. It was first stated by Benoît Paul Émile Clapeyron in 1834 as a combination of the empirical Boyle's law, Charles's law, Avogadro's law, and Gay-Lussac's law.

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